Course Structure

Chemistry

(A)Major Core Courses

Sl. No.	Sem.	Type of Course	Name of Course	Credits	Marks
1.	I	MJC-1 (T)	Inorganic Chemistry : Atomic Structure & Chemical Bonding Organic Chemistry : Fundamental of Organic Chemistry, stereochemistry & aliphatic hydrocarbons	4	100
		MJC-1 (P)	Inorganic Chemistry Lab: Volumetric analysis. Organic Chemistry Lab : Detection, Purification, Separation of organic Compouds.	2	100
2.	II	MJC-2 (T)	Physical Chemistry : States of matter and ionic equilibria.	4	100
		MJC-2 (P)	Physical Chemistry Lab: Determination of Physical Properties of liquids and pH- metry.	2	100
3.	III	MJC-3	Organic Chemistry : Basics and Hydrocarbons	5	100
4.	III	MJC-4 (T)	Physical Chemistry II: Chemical Thermodynamics and its Applications (T)	3	100
		MJC-4 (P)	Physical Chemistry II: Chemical Thermodynamics and its Applications (P)	1	100
5.	IV	MJC-5 (T)	Inorganic Chemistry II: s- and p Block Elements (T)	3	100
		MJC-5 (P)	Inorganic Chemistry II: s- and p Block Elements (P)	2	100
6.	IV	MJC-6 (T)	Organic Chemistry II: Oxygen Containing Functional Groups (T)	3	100
		MJC-6 (P)	Organic Chemistry II: Oxygen Containing Functional Groups (P)	2	100
7.	IV	MJC-7	Physical Chemistry III: Phase Equilibria and Electrochemical Cells	5	100
8.	V	MJC-8 (T)	Inorganic Chemistry III: Coordination Chemistry (T)	3	100
		MJC-8 (P)	Inorganic Chemistry III: Coordination Chemistry (P)	2	100
9.	V	MJC-9	Organic Chemistry III: Heterocyclic Chemistry	5	100
10.	VI	MJC-10 (T)	Physical Chemistry IV: Conductance & Chemical Kinetics (T)	3	100
		MJC-10 (P)	Physical Chemistry IV: Conductance & Chemical Kinetics (P)	1	100
11.	VI	MJC-11 (T)	Organic Chemistry IV: Biomolecules (T)	3	100
		MJC-11 (P)	Organic Chemistry IV: Biomolecules (P)	2	100
12.	VI	MJC-12	Physical Chemistry V: Quantum Chemistry & Spectroscopy	5	100
13.	VII	MJC-13	Inorganic Chemistry IV: Organometallic Chemistry	5	100
14.	VII	MJC-14	Research Methodology	5	100
15.	VII	MJC-15	Organic Chemistry V: Spectroscopy	6	100
16.	VIII	MJC-16	Analytical Methods in Chemistry (T)	4	100

Sub Total = 80

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Chemistry

SEMESTER – I

MJC-1(T) : Inorganic Chemistry: Atomic Structure, Chemical Bonding and fundamentals of organic Chemistry

Course Objective

The Objective of CBCS based four year undergraduate Programme (FYUGP) in Chemistry Hons for Semester I & II, Specially for Major & Minor course is to provide the clear conception and understanding about theory and practical course mentioned in the syllabus.

Course Outcomes

After the completion of the course, the students will be able to understand the following:

: the model of an atom including the related various principles.

: the principles of bonding as well as shapes and structure of covalent molecules.

: Intial step of research in Organic Chemistry viz-Detectionl, Seperation and Purfication of Organic Compounds.

MJC-1(T)

(Theory: 4 credits)

Inorganic Chemistry : Atomic Structure and Chemical Bonding		
Unit	**Topics to be covered**	**No. of Lectures**
1	**Atomic Structure :** Bohr's Theory, its limitations and atomic spectrum of hydrogen atom. Wave mechanics: de Broglie equation, Heisenberg's Uncertainty Principle and its significance, Schrodinger's wave equation, significance of wave function. Quantum Numbers and their significance. Normalized and orthogonal wave functions. Sign of wave functions. Radial and angular wave functions for hydrogen atom. Radial and angular distribution curves. Shapes of s, p, d and f orbitals. Contour boundary and probability diagrams. Pauli's Exclusions Principle, Hund's Rule of maximum multiplicity, Aufbau's principle and its limitations, Variations of orbital energy with atomic number.	10
2	**Chemical Bonding:** (i) Ionic bond: General characteristics, types of ions size effects, radius ratio rule and its limitations. Packing of ions in crystals. Born-Lande equation with derivation and importance of Kapustinskii expression for lattice energy. Madelung constant, Born-Haber cycle and its application, Solvation energy. (ii) Covalent bond: Lewis structure, Valence Bond Theory (Heitler-London approach). Energetics of hybridization, equivalent and non-equivalent hybrid orbitals. Bent's rule, Resonance and resonance energy, Molecular orbital theory. Molecular orbital diagrams of diatomic and simple polyatomic molecules N_2, O_2,C_2, B_2, F_2 & CO, NO and their ions; hydrogen chloride, berrylium fluoride, carbon dioxide, (idea of s-p mixing and orbital interaction to be given). Formal charge, Valence shell electron pair repulsion theory (VSEPR), shapes of simple molecules and ions containing lone pairs and bond pairs of electrons, multiple bonding (sigma and pi bond approach) and bond lengths. Covalent character in ionic compounds, polarizing power and polaizability. Fajan's rule and consequences of polarization. Ionic character in covalent compounds: Bond moment, dipole moment and electronegativity difference. (iii) Metallic bond: Qualitative idea of valence bond and band theories. Semiconductors and insulators, defects in solids. (iv) Weak Chemical Forces: van der Waals forces, ion-dipole forces, dipole-dipole interactions, induced dipole interactions, Instantaneous dipole-induced dipole interactions. Repulsive forces, Hydrogen Bonding (theories of hydrogen bonding, valence bond treatment) Effects of chemical force, melting and boiling points, solubility energetics of dissolution process.	10

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	Organic Chemistry:-Fundamentals, Sterochemistry and hydrocarbons.	
3	**Fundamentals of Organic Chemistry:** Physical Effects, Electronic Displacements: Inductive Effect, Electromeric Effect, Resonance and Hyperconjugation. Cleavage of Bonds: Homolysis and Heterolysis. Structure, shape and reactivity of organic molecules, Nucleophiles and electrophiles. Reactive Intermediates: Carbocations, Carbanions and free radicals. Strength of organic acids and bases: Comparative study with emphasis on factors affecting pK values. Aromaticity: Benzenoids and Hückel's rule.	05
4	**Stereochemistry** Conformations with respect to ethane, butane and cyclohexane. Interconversion of Wedge Formula, Newmann, Sawhorse and Fischer representations. Concept of chirality (upto two carbon atoms). Configuration: Geometrical and Optical isomerism; Enantiomerism, Diastereomerism and Meso compounds). Threo and erythro; D and L; *cis - trans* nomenclature; CIP Rules: R/ S (for upto 2 chiral carbon atoms) and E / Z Nomenclature (for upto two C=C systems).	05
5	**Aliphatic Hydrocarbons** Functional group approach for the following reactions (preparations & reactions) to be studied in context to their structure. **Alkanes:** (Upto 5 Carbons). *Preparation:* Catalytic hydrogenation, Wurtz reaction, Kolbe's synthesis, from Grignard reagent. *Reactions:* Free radical Substitution: Halogenation. **Alkenes:** (Upto 5 Carbons) *Preparation:* Elimination reactions: Dehydration of alkenes and dehydrohalogenation of alkyl halides (Saytzeff's rule); cis alkenes (Partial catalytic hydrogenation) and trans alkenes (Birch reduction). *Reactions:* cis-addition (alk. KMnO4) and trans-addition (bromine), Addition of HX (Markownikoff's and anti-Markownikoff's addition), Hydration, Ozonolysis, oxymecuration-demercuration, Hydroboration-oxidation. **Alkynes**: (Upto 5 Carbons) *Preparation:* Acetylene from CaC_2 and conversion into higher alkynes; by dehalogenation of tetra halides and dehydrohalogenation of vicinal-dihalides. *Reactions:* formation of metal acetylides, addition of bromine and alkaline $KMnO_4$, ozonolysis and oxidation with hot alk. $KMnO_4$.	10
	TOTAL	40

Suggested Readings :

1. Advanced Inorganic Chemistry, F.A. Cotton, G. Wilkinson.
2. Concise Inorganic Chemistry, J.D. Lee, Blackwell Science, 2001.
3. Inorganic Chemistry, J.E. Huheey, E.A. Keiter and R.I. Keiter, Pearson Education Asia, 2000.
4. Inorganic Chemistry, ELBS 2nd Edition, D.F. Shriver, P.W. Atkins and C.H. Langford. Oxford University Press 2002.
5. Principles of Inorganic Chemistry. B.R. Puri, L.R. Sharma, Jauhar S.P., S.N. Chand & Co.
6. Inorganic Chemistry, 3rd Edition (ISE) A.G. Sharpe Addison Wesley.
7. Organic Chemisry - Graham Solomons
8. Sterochemistry – Conformation and Mechanism : P.S Kalsi
9. Organic Chemistry - Morrison & Boyd

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MJC -1 (P) : 2 Credits

Inorganic and Organic Chemistry Practical

Course Outcomes

After finishing this Practical Course, Studens will be skilled in-

1. Understanding and preparing solution of different strengths.
2. Calculating the neutral point in different titrations.
3. Initial research steps involved in Organic Chemistry.

Practical- 1. Inorganic Chemisry Practical

a. Acidimetry and Alkalimetry
b. Preparation and dilution of standard solutions.
c. Permangnatometry / dichromatry
d. Iodometry / iodimetry

Practical- 2.Organic Chemistry Practical

Detection of elements, separation and purification of Organic Compounds and Hydrolysis of ester.

Suggested Readings :

1. Practical inorganic chemistry : Shikha Gulati and J. L . Sharma
2. Practical Chemistry : Dr O .P. Pandey , D.N. Bajpayi & Giri.
3. Quantitative Chemical analaysis: A.I. Vogel, Prentice Hall Publication.
4. Text book of practical Organic Chemistry: A.I. Vogal, Prentice Hall Publication.
5. Practical Organic Chemistry, F.G. Mann & B.C. Saunders,Orient long man.

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